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Objectives. After completing this section, you should be able to write the equilibrium constant expression for a given reaction. assess, qualitatively, how far a reaction will proceed in a given direction, given the value of K_{eq} ; explain the difference between rate and equilibrium.

6.8: Describing a Reaction - Equilibria, Rates, and Energy ...

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developed as a correlation of reactivities (rates) and equilibria in reactions of meta- and para-substituted benzene derivatives. As early as 1912 it was noted that in reactions of m- and p-substituted benzene derivatives there were regularities in substituent effects. For example, whatever effect a p-NO₂ group

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equilibrium constants with changes in concentration, pressure or in the presence of a catalyst

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