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Making Sense of Chemical Structures

Basic Chemistry for Biology Part 4:

Covalent Bonding and Structural Formulas

Understanding Atoms, elements, and

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*molecules Part #1 (9min) Chapter 3 The
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1 of 2 Chapter 2 — Chemistry Basic*

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#1 (Lt-02) Atomic Structure Part-2 //

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Ch. 2 Essential Questions: Must be answered on a separate sheet of paper

- 1) Differentiate between matter and energy.
- 2) Explain how each of the 4 types of energy are different.
- 3) Draw an atom of Carbon and label its Protons, Neutrons, and Electrons.
- 4) Name the 4 most abundant elements found in the Human

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Body. 5) Compare and contrast Ionic and Covalent Bonds.

~~Chapter 2: Basic Chemistry Name: Period~~

Figure 2—1 is a diagram of an atom. Select two different colors and use them to color the coding circles and corresponding structures on the figure. Complete this

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exercise by responding to the questions that follow, referring to the atom in this figure. Insert your answers in the answer blanks provided. O Nucleus Electrons 1. 2. 3. 5. Figure ...

~~Mr. Burghardt Ponderosa High School
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The Muscular System. Chapter 7 The Nervous System.

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Basic Chemistry Although it is possible to study anatomy without much reference to chemistry, chemical reactions underlie all body processes-movement, digestion, the pumping of your heart, and even your thoughts. Chapter 2 Basic Chemistry - studylib.net specifically get lead by on-line.

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Chapter 2 Basic Chemistry. Objectives. Differentiate clearly between matter and energy. List the major energy forms, and provide one example of how each energy form is used in the body. List four...

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~~Chapter 2 Basic Chemistry – Anatomy and Physiology~~

Basic Chemistry Concepts of Matter and Energy Composition of Matter Molecules and Compounds Chemical Bonds and Chemical Reactions Biochemistry.
Concepts of Matter and Energy. Matter.

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Matter is anything that occupies space and has weight. Matter exist in three states: gas, liquid, and solid.

~~Chapter 2: Basic Chemistry~~

Anatomy and Biology Chemistry Review

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1 Chapter 2.1, 2.2 Review Packet –

Answer Key 2.1 The Nature of Matter

Lesson Summary Atoms The atom is the
basic unit of matter, made up of three
subatomic particles. Protons have a

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positive charge and neutrons carry no charge. Strong forces bind protons and neutrons together in the nucleus. An electron is a negatively charged particle that has only about $1/1840$ the mass of a proton.

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