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Nuclear Chemistry Section

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Chemistry Notes (answers)

Chapter 21: Nuclear

Chemistry 21.1: The Nature

of Nuclear Reactions

Nucleons: - the particles

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Section 3 Answers

that make up a nucleus of an atom (protons, $(1\ 1\ p\ +\ or\ 1\ 1\ H)$ and neutrons, $(1\ 0\ n)$).
Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

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Chapter 21 Nuclear Chemistry Review Answers

Nuclear Chemistry Nuclear
Transformations • Rutherford
in 1919 performed the first
nuclear transformation. •
The transmutations are
sometimes represented by

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Section 3 Answers

listing in order, the target nucleus, the bombarding particle, the ejecting particle and the product nucleus. • The above equation becomes:

$${}_{7}^{14}\text{N} + {}_{2}^{4}\text{He} \rightarrow {}_{8}^{17}\text{O} + {}_{1}^{1}\text{H}$$

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Chapter 21 Nuclear Chemistry

Nuclear chemistry is the study of reactions that involve changes in nuclear structure. The chapter on atoms, molecules, and ions introduced the basic idea of nuclear structure, that the

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nucleus of an atom is

composed of protons and,

with the exception of ${}^1_1\text{H}$,

${}^1_0\text{n}$, neutrons. Recall that

the number of protons in the

nucleus is called the atomic

number (Z) of the element,

and the sum of the number of

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protons and the number of
neutrons is the mass number
(A) .

21.1 Nuclear Structure and Stability - Chemistry

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Chemistry. Modern Chemistry

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175 Nuclear Chemistry CHAPTER
21 REVIEW Nuclear Chemistry
SECTION 4 SHORT ANSWER

Answer the following questions in the space provided. 1. Match each of the following statements with the process(es) to

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which they apply, using one
of the choices below: (1)
fission only (3) both
fission and fusion

Chapter 21 Nuclear Chemistry **Section 3 Answers**

Chapter 21 Nuclear

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Chemistry. Section 21.2

Nuclear Reactions and
Energy. Objectives: Compare
and Contrast Nuclear fission
and Fusion, Demonstrate
Equations that Represent the
Changes that Occur During
Radioactive Decay, Trace the

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Section 3 Answers
Operation and Structure of a
Nuclear Reactor. The Power
of the Nucleus. Nuclear
reactions involve enormous
energy changes.

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Section 3 Answers
By the end of this section,
you will be able to:

Identify common particles
and energies involved in
nuclear reactions Write and
balance nuclear equations
Changes of nuclei that
result in changes in their

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atomic numbers, mass numbers, or energy states are nuclear reactions.

21.2 Nuclear Equations - Chemistry

Chapter 21 Nuclear
Chemistry. STUDY. PLAY.

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nuclear reaction. a change in the composition or makeup of the nucleus. nuclear decay. the breakdown of nuclei by giving off particles or electromagnetic radiation, natural. fission. splitting of an atom into

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smaller atoms; requires lots
of energy, not at high temp,
chain reaction.

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Answer Key

Nuclear reactions usually change one type of nucleus into another; chemical changes rearrange atoms.

Nuclear reactions involve

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Section 3 Answers
much larger energies than
chemical reactions and have
measurable mass changes. 9.
(a), (b), (c), (d), and (e)

Answer Key Chapter 21 - Chemistry 2e | OpenStax

Chapter 21. Nuclear

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Chemistry. 21.5 Uses of Radioisotopes. Learning Objectives. By the end of this section, you will be able to: List common applications of radioactive isotopes; Radioactive isotopes have the same

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Section 3 Answers as
stable isotopes of the same
element, but they emit
radiation, which can be
detected. If we replace one
(or ...

21.5 Uses of Radioisotopes -

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Section 3 Answers

If the parent nuclide undergoing β^- decay lies below the band of stability (refer to Chapter 21.1 Nuclear Structure and Stability), the daughter nuclide will lie closer to

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Section 3 Answers (?) decay is the emission of an electron from a nucleus. Iodine-131 is an example of a nuclide that undergoes β^- decay:

21.3 Radioactive Decay - Chemistry

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Questions and Exercises -

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Scott R.; Ball, David W.,
ISBN-10: 0534420125,
ISBN-13: 978-0-53442-012-3,
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692 Chapter 3 Answers Nuclear
Chemistry 16.1 The Nucleus
and Radioactivity Our
journey into the center of
the atom begins with a brief
review. You learned in
Chapter 3 that the protons
and neutrons in each atom

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are found in a tiny, central nucleus that measures about $1/100,000$ the diameter of the atom itself. You also learned

Chapter 16 Nuclear Chemistry

Chemistry and Chemical

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Reactivity (9th Edition)

answers to Chapter 6 The
Structure of Atoms - 6-2

Quantization: Planck,

Einstein, Energy, and

Photons - Review & Check for

Section 6-2 - Page 226 2

including work step by step

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